

10/643,090

EAST Search History

Ref #	Hits	Search Query	DBs	Default Operator	Plurals	Time Stamp
L1	255	(556/110).CCLS.	US-PGPUB; USPAT; EPO; JPO	OR	OFF	2007/02/12 18:43

10/643,090

(FILE 'HOME' ENTERED AT 14:24:30 ON 12 FEB 2007)

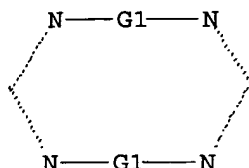
FILE 'REGISTRY' ENTERED AT 14:24:40 ON 12 FEB 2007

L1 STRUCTURE UPLOADED

=> d l1

L1 HAS NO ANSWERS

L1 STR



G1 Ag,Au,Cu

Structure attributes must be viewed using STN Express query preparation.

=> s l1 full

FULL SEARCH INITIATED 14:25:10 FILE 'REGISTRY'

FULL SCREEN SEARCH COMPLETED - 37064 TO ITERATE

100.0% PROCESSED 37064 ITERATIONS
SEARCH TIME: 00.00.01

809 ANSWERS

L2 809 SEA SSS FUL L1

=> fil caplus

COST IN U.S. DOLLARS

SINCE FILE	TOTAL
ENTRY	SESSION
172.10	172.31

FULL ESTIMATED COST

FILE 'CAPLUS' ENTERED AT 14:25:15 ON 12 FEB 2007

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FILE LAST UPDATED: 11 Feb 2007 (20070211/ED)

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=> s l2

L3 303 L2

=> s 13 and py<2003
 22869379 PY<2003
 L4 228 L3 AND PY<2003

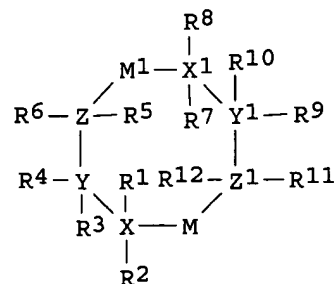
=> s 14 and thin films
 593662 THIN
 811709 FILMS
 149945 THIN FILMS
 (THIN(W) FILMS)
 L5 0 L4 AND THIN FILMS

=> s 14 and CVD
 70280 CVD
 L6 1 L4 AND CVD

=> d bib abs

L6 ANSWER 1 OF 1 CAPLUS COPYRIGHT 2007 ACS on STN
 AN 2001:747167 CAPLUS
 DN 135:273081
 TI Preparation of metal volatile precursors for deposition of metals and
 metal-containing films
 IN Morman, John Anthony Thomas; Roberts, David Allen; Farnia, Morteza
 PA Air Products and Chemicals, Inc., USA
 SO Eur. Pat. Appl., 21 pp.
 CODEN: EPXXDW
 DT Patent
 LA English
 FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 1142894	A2	20011010	EP 2001-108053	20010329 <--
	EP 1142894	A3	20030423		
	EP 1142894	B1	20050112		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	US 2002013487	A1	20020131	US 2001-791409	20010222 <--
	TW 490502	B	20020611	TW 2001-90107555	20010329 <--
	AT 286902	T	20050115	AT 2001-108053	20010329
	JP 2002069088	A	20020308	JP 2001-104544	20010403 <--
	JP 3593051	B2	20041124		
	HK 1039944	A1	20050805	HK 2002-101237	20020220
PRAI	US 2000-194285P	P	20000403		
	US 2001-791409	A	20010222		
OS	CASREACT 135:273081; MARPAT 135:273081				
GI					



I

AB This invention is directed to a group of novel homologous eight membered ring compds. I having a metal, such as copper, reversibly bound in the ring and containing carbon, nitrogen, silicon and/or other metals. A

structural representation of the compds. I (M, M' = Cu, Ag, Au, Ir; X, X' = N, O; Y, Y' = Si, C, Sn, Ge, B; Z, Z' = C, N, O; substituents represented by R1, R2, R3, R4, R5, R6, R7, R8, R9, R10, R11, R12 will vary depending on the ring atom to which they are attached). This invention is also directed to depositing metal and metal-containing films on a substrate, under ALD or CVD conditions, using the above novel compds. as precursors. Thus, reaction of dimethylaminochloromethyldimethylsilane with Mg in THF followed by treatment with cuprous chloride gave [-CuNMe2SiMe2CH2CuNMe2SiMe2CH2-].

=> s 14 and copper

923086 COPPER

L7 188 L4 AND COPPER

=> s 17 and amidinate

326 AMIDINATE

L8 1 L7 AND AMIDINATE

=> d bib abs

L8 ANSWER 1 OF 1 CAPLUS COPYRIGHT 2007 ACS on STN

AN 1997:314951 CAPLUS

DN 127:5420

TI Living Polymerization of Carbodiimides Initiated by Copper(I) and Copper(II) Amidinate Complexes

AU Shibayama, Koichi; Seidel, Scott W.; Novak, Bruce M.

CS Department of Polymer Science and Engineering, University of Massachusetts, Amherst, MA, 01003, USA

SO Macromolecules (1997), 30(11), 3159-3163

CODEN: MAMOBX; ISSN: 0024-9297

PB American Chemical Society

DT Journal

LA English

AB Robust catalysts based on copper(I) and copper(II) amidinate complexes initiate living polymerization of carbodiimide. The tolerance of these complexes to impurities is illustrated by the fact that they cleanly initiate the polymerization of carbodiimides under air and oxygen. They are even active in the presence of water, but both mol. wts. and yields tend to be lower than in dry solvents. The catalytic activity of a copper(II) amidinato complex is almost equal that of reported titanium(IV) initiators. Both oxidation states are active, but Cu(II) complexes are more active in terms of rates of reaction. Regardless of the oxidation state of the initial complex, all polymns. run in the presence of oxygen proceed through the Cu(II) oxidation state. Mechanistic studies indicate that the carbodiimides insert into one of the copper-amidinate bonds, thus becoming the end group of the growing polymer chain. The resultant polycarbodiimides from bulk polymns. were isolated, after dissolving to toluene, by precipitation into excess methanol,

and

lyophilization from benzene, as a spongy white solid. Anal. of these systems by gel permeation chromatog.-light scattering measurements (GPC-LS) and preliminary kinetic anal. suggest this system to be living. Polycarbodiimides adopt extended-chain, helical conformations; data from X-ray scattering studies and mol. modeling indicate that polycarbodiimides display a 6/1 helix in the solid state, and viscometry and light scattering data indicate that this extended-chain conformation persists in solution

RE.CNT 22 THERE ARE 22 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

=> s 14 and gold

240181 GOLD

L9 8 L4 AND GOLD

=> d 1-8 bib abs

L9 ANSWER 1 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN
AN 2000:119017 CAPLUS
DN 132:245405
TI One-dimensional metal string complexes
AU Peng, Shie-Ming; Wang, Chih-Chieh; Jang, Yih-Lih; Chen, Yu-Hua; Li, Feng-Yin; Mou, Chung-Yuan; Leung, Man-Kit
CS Department of Chemistry, National Taiwan University, Taipei, 106, Taiwan
SO Journal of Magnetism and Magnetic Materials (2000), 209, 80-83
CODEN: JMMMDL; ISSN: 0304-8853
PB Elsevier Science B.V.
DT Journal
LA English
AB Metal string complexes with different metal ions and various lengths were successfully synthesized and structurally characterized. The crystal and mol. structure of nonanickel complex $[\text{Ni}_9(\mu_9\text{-peptea})_4\text{Cl}_2]$ which is the longest metal string at this moment is reported. Crystal data for $[\text{Ni}_9(\mu_9\text{-peptea})_4\text{Cl}_2] \cdot (\text{C}_2\text{H}_4\text{Cl}_2)_{10}$: monoclinic, space group $\text{C}2/c$, a 27.0453(6), b 42.6678(5), c 37.4268(6) Å, β 109.539(1)°, V = 40702(1) Å³, Z = 12, R_F = 12% for the observed reflections ($I > 2\sigma(I)$). The band structure calcn. of infinite 1-dimensional metal chains of NiII and CrII based on the structure of nonanickel complex will be presented. Finally, the metal string complexes $[\text{MnII}(\text{L}_4\text{X}_2)]$ will be chemisorbed on the Ag or Au surface and the potential application of these complexes as a mol. metal wire will be demonstrated.

RE.CNT 13 THERE ARE 13 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L9 ANSWER 2 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN
AN 1995:278857 CAPLUS
DN 122:95255
TI A luminescent heterometallic $\text{Au} \cdots \text{Cu}$ complex.
Spectroscopic properties and crystal structures of $[\text{Au}(\text{PPh}_3)(\text{C}_7\text{H}_5\text{N}_2)]$ and $[\{\text{Au}(\text{PPh}_3)(\mu\text{-C}_7\text{H}_5\text{N}_2)\text{Cu}(\mu\text{-C}_7\text{H}_5\text{N}_2)\}_2]$ ($\text{C}_7\text{H}_5\text{N}_2$ = 7 azaindolate)
AU Chan, Chi-Keung; Guo, Chun-Xiao; Cheung, Kung-Kai; Li, Dan; Che, Chi-Ming
CS Dep. Chem., Univ. Hong Kong, Hong Kong, Hong Kong
SO Journal of the Chemical Society, Dalton Transactions: Inorganic Chemistry (1994), (24), 3677-82
CODEN: JCDTBI; ISSN: 0300-9246
PB Royal Society of Chemistry
DT Journal
LA English
AB The mono- and tetra-nuclear complexes $[\text{Au}(\text{PPh}_3)(\text{C}_7\text{H}_5\text{N}_2)]$ (1) and $[\{\text{Au}(\text{PPh}_3)(\mu\text{-C}_7\text{H}_5\text{N}_2)\text{Cu}(\mu\text{-C}_7\text{H}_5\text{N}_2)\}_2]$ (2), the latter containing 7-azaindolate ($\text{C}_7\text{H}_5\text{N}_2$ -) bridging ligands, have been prepared and their crystal structures determined. The intramol. $\text{Au} \cdots \text{Cu}$ and $\text{Cu} \cdots \text{Cu}$ sepns. in 2 are 3.0104(6) and 2.941(1) Å resp. Extended-Hueckel mol.-orbital calcns. revealed a weak $\text{Au} \cdots \text{Cu}$ bonding interaction but on $\text{Cu} \cdots \text{Cu}$ interaction. In MeCN, both complexes display intense intraligand emission at 510 nm upon photoexcitation with UV/visible light at room temperature. In the solid state 2 shows an emission at 550 nm.

L9 ANSWER 3 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN
AN 1991:54777 CAPLUS
DN 114:54777
TI $\text{Ag}_2[\text{PhC}(\text{NSiMe}_3)_2]_2$ and $\text{Au}_2[\text{PhC}(\text{NSiMe}_3)_2]_2$: amidinato complexes with short metal-metal distances
AU Fenske, Dieter; Baum, Gerhard; Zinn, Alfred; Dehnicke, Kurt
CS Inst. Anorg. Chem., Univ. Karlsruhe, Karlsruhe, D-7500, Germany

SO Zeitschrift fuer Naturforschung, B: Chemical Sciences (1990),
45(9), 1273-8
CODEN: ZNBSEN; ISSN: 0932-0776

DT Journal

LA German

AB $M_2[PhC(NSiMe_3)_2]_2$ ($M = Ag, Au$) were prepared by the reaction of $PhC(NSiMe_3)[N(SiMe_3)_2]$ with AgO_2CCH_3 , and $ClAuCO$, resp., in MeCN suspensions. Both complexes were characterized by IR spectroscopy and X-ray structure detns.: triclinic, space group $P\bar{1}$, $Z = 2$.
 $Ag_2[PhC(NSiMe_3)_2]_2$: a 1132.5(5), b 1221.4(6), c 1407.9(8) pm, α 107.85(3), β 106.63(4), γ 91.54(3)°, $R = 0.023$.
 $Au_2[PhC(NSiMe_3)_2]_2$: a 1119.6(6), b 1219.0(8), c 1391.7(9) pm, α 108.74(4), β 106.70(5), γ 91.49(4)°; $R = 0.04$. The compds., which are isotypical, form dimeric mols. with short metal-metal contacts of $Ag...Ag = 265.5(2)$ and $Au...Au = 264.6(2)$ pm and 2 chelating amidinato ligands with $Ag-N$ bond lengths of (averaged) 212.4(2) pm and $Au-N$ bond lengths of 206.0(7) pm.

L9 ANSWER 4 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN

AN 1989:515296 CAPLUS

DN 111:115296

TI N,N,N'-Tris(trimethylsilyl)organoamidinate as reagents in complex chemistry

AU Dehnicke, Kurt; Ergezinger, Christina; Hartmann, Eva; Zinn, Alfred; Hoesler, Klaus

CS Fachbereich Chem., Univ. Marburg, Marburg, D-3550, Fed. Rep. Ger.

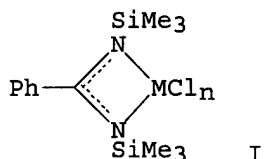
SO Journal of Organometallic Chemistry (1988), 352(1-2), C1-C4
CODEN: JORCAI; ISSN: 0022-328X

DT Journal

LA German

OS CASREACT 111:115296

GI



AB N,N,N'-Tris(trimethylsilyl)phenylamidines, $PhC(NSiMe_3)[N(SiMe_3)_2]$, react with many main group element and transition metal halides to give mainly organoamidinato complexes I (e.g., $M = Al, Ga, In, Sn, Sb, Bi, Se, Te, Ti, Zr, Hf, Nb, Ta, W, Cu, Au$, $n = 2, 3, 4$).

L9 ANSWER 5 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN

AN 1987:598598 CAPLUS

DN 107:198598

TI η^3 -Allylrhodium complexes with azolate ligands. Crystal structures of the trinuclear 1,2,4-triazolate complexes $tris[di-\eta\text{-allyl-}\mu\text{-(1,2,4-triazolato-N}_2\text{N}_4\text{)-rhodium}]$ and $1,1\text{-di-}\eta\text{-allyl-2,2,3,3-tetracarbonyl-1,2-}\mu\text{-chloro-3-chloro-}\mu^3\text{-1,2,4-triazolato-N}_1\text{N}_2\text{N}_4\text{-trirhodium-1,2-dichloroethane (2/1)}$

AU Oro, Luis A.; Pinillos, M. Teresa; Tejel, Cristina; Foces-Foces, Concha; Cano, Felix H.

CS Inst. Cienc. Mater. Aragon, Univ. Zaragoza, Zaragoza, 50009, Spain

SO Journal of the Chemical Society, Dalton Transactions: Inorganic Chemistry (1972-1999) (1986), (10), 2193-200
CODEN: JCDTBI; ISSN: 0300-9246

DT Journal

LA English
OS CASREACT 107:198598
GI

* STRUCTURE DIAGRAM TOO LARGE FOR DISPLAY - AVAILABLE VIA OFFLINE PRINT *

AB $\text{Rh}_3(\mu\text{-L})_3(\eta^3\text{-C}_3\text{H}_5)_6$ [L = tetrazolate (I; ttz), 1,2,4-triazolate (II; tz), benzotriazolate, or imidazolate] were prepared by reaction of $\text{Rh}(\text{acac})(\eta^3\text{-C}_3\text{H}_5)_2$ (acac = acetylacetonate) with azoles HL. A single crystal x-ray diffraction study on complex II showed trinuclear complexes in which each rhodium atom is bridged by the 1,2,4-triazolate ligands through the nonadjacent nitrogen atoms to form a multiat. ring. Complexes I and II react with $[\text{Rh}_2(\mu\text{-Cl})_2(\text{CO})_4]$ to give trinuclear complexes $[\text{Rh}_3(\mu_3\text{-L})(\mu\text{-Cl})\text{Cl}(\eta^3\text{-C}_3\text{H}_5)_2(\text{CO})_4]$ [L = tz (III) or ttz]. The crystal structure of complex III showed trinuclear mols. in which a chlorine atom and a triazolate bridge $\text{Rh}(\text{CO})_2$ and $\text{Rh}(\eta^3\text{-C}_3\text{H}_5)_2$ moieties; in addition the remaining nitrogen atom of the triazolate ligand completes the square planar coordination of a $\text{RhCl}(\text{CO})_2$ unit. Complex II reacts with $[\text{Rh}_2(\mu\text{-Cl})_2(\eta^3\text{-C}_3\text{H}_5)_4]$ (IV) to give the trinuclear complex $[\text{Rh}_3(\mu_3\text{-tz})(\mu\text{-Cl})\text{Cl}(\eta^3\text{-C}_3\text{H}_5)_6]$. Heterobridged binuclear pyrazolate (pz) complexes of formula $[\text{Rh}_2(\mu\text{-pz})(\mu\text{-X})(\eta^3\text{-C}_3\text{H}_5)_4]$ (X = Cl, OH, or N3) have also been prepared from IV and pz followed by addition of KOH or NaN_3 .

L9 ANSWER 6 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN

AN 1984:630691 CAPLUS

DN 101:230691

TI A new approach to the synthesis of cationic, neutral or anionic diarylgold(III) complexes

AU Uson, R.; Laguna, A.; Arrese, M. L.

CS Dep. Inorg. Chem., Univ. Zaragoza, Spain

SO Synthesis and Reactivity in Inorganic and Metal-Organic Chemistry (1984), 14(4), 557-67

CODEN: SRIMCN; ISSN: 0094-5714

DT Journal

LA English

AB The reaction of $[\text{Au}(\text{C}_6\text{F}_5)_2\text{Cl}]_2$ or $\text{Bu}_4\text{N}[\text{Au}(\text{C}_6\text{F}_5)_2\text{Cl}_2]$ with AgClO_4 renders stable solns. (in acetone or Et_2O) of $[\text{Au}(\text{C}_6\text{F}_5)_2\text{S}_2]^+$ (S = AcMe, Et_2O) which react (a) with neutral ligands to give cationic complexes of the type $[\text{Au}(\text{C}_6\text{F}_5)_2\text{L}_2]^+$ (L = tetrahydrothiophene, PPh_2Me ; L2 = o-phenylenebis(dimethylarsine), bipy); b) with anionic ligands to give the neutral complexes $\text{Au}(\text{C}_6\text{F}_5)_2(\text{acac})$ of $[\text{Au}(\text{C}_6\text{F}_5)_2]_2(\text{BiBzIm})$ (BiBzIm = bibenzimidazolate) or the anionic ones $[\text{Au}(\text{C}_6\text{F}_5)_2(\text{CN})_2]^-$. Treatment of the latter with AgClO_4 leads to the tetrameric neutral complex $[\text{Au}(\text{C}_6\text{F}_5)_2(\text{CN})_2\text{Ag}]_4$.

L9 ANSWER 7 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN

AN 1981:631085 CAPLUS

DN 95:231085

TI Bibenzimidazolate as bridging ligand in bi-, tri- and tetranuclear gold(I) or gold(I) and palladium(II) complexes

AU Uson, R.; Gimeno, J.; Fornies, J.; Martinez, F.; Fernandez, C.

CS Dep. Inorg. Chem., Univ. Saragossa, Saragossa, Spain

SO Inorganica Chimica Acta (1981), 54(3), L95-L96

CODEN: ICHAA3; ISSN: 0020-1693

DT Journal

LA English

AB K or Tl salts of 2,2'-bibenzimidazole (H_2Q) in $\text{CH}_2\text{Cl}_2\text{-MeOH}$ suspension react with $\text{AuCl}(\text{PPh}_3)$ and AuCl_2L [L = bis(diphenylphosphino)methane] to give $(\mu\text{-Q})[\text{AuPPh}_3]_2$ and $(\mu\text{-Q})\text{Au}_2\text{L}$, resp. AuClL' (L' = py, tetrahydrothiophene) react with Q2- to give $[\text{Au}_2(\mu\text{-Q})]_x$.

(μ -Q) [AuPPh₃]₂ reacts with Au(ClO₄)PPh₃ to give {(μ -Q) [AuPPh₃]₃}ClO₄. (μ -Q)Au₂L reacts with trans-PdCl₂L'₂ (L' = tetrahydrothiophene) to give Cl₂Pd(μ -Q) [Au₂L]. (μ -Q)Au₂L reacts with Au(ClO₄)PPh₃ in a 1:2 ratio or with (μ -L) [Au(ClO₄)]₂ in a 1:1 ratio to give {[Ph₃PAu]₂ μ -Q[Au₂L]} and { μ -Q[Au₂L]₂}(ClO₄)₂. (μ -Q)Au₂L reacts with (μ -Cl)₂Pd₂L'₂(C₆F₅)₂ (L' = tetrahydrothiophene) or [(μ -Br)₂Pd₂(C₆F₅)₄]₂- to give Cl(C₆F₅)Pd(μ -Q)Au₂L or (C₆F₅)₂Pd(μ -Q)Au₂L, resp.

L9 ANSWER 8 OF 8 CAPLUS COPYRIGHT 2007 ACS on STN
 AN 1981:533062 CAPLUS
 DN 95:133062
 TI New ways for the preparation of heterocyclic gold(III) complexes
 AU Uson, R.; Vicente, J.; Chicote, M. T.
 CS Dep. Inorg. Chem., Univ. Zaragoza, Spain
 SO Journal of Organometallic Chemistry (1981), 209(2), 271-9
 CODEN: JORCAI; ISSN: 0022-328X
 DT Journal
 LA English
 AB Treating 1,1-dimethyl-2,3,4,5-tetraphenylstannacyclopentadiene [(Ph₄C₄)SnMe₂] with AuCl₃(Q) (Q = tetrahydrothiophene) gave auracyclopentadiene complexes [(Ph₄C₄)AuCl(Q)] or [(Ph₄C₄)AuCl]₂, which react with Tl(acac) (acac = acetylacetonato) to give [(Ph₄C₄)Au(acac)]. Reaction of the last with protonated ligands gave neutral mononuclear or binuclear complexes [(Ph₄C₄)Au(XL)] (XL is a monoanionic bidentate ligand) or [(Ph₄C₄)Au(X₂L₂)Au(Ph₄C₄)] (X₂L₂ is a dianionic tetradentate ligand). Treating the acetylacetonate complex with ammonium or pyridinium salts gave mononuclear neutral complexes [(Ph₄C₄)Au(X)(L)] (X = AcO, CF₃CO₂; L = NH₃, py) or the cationic complexes [(Ph₄C₄)AuLL₁]⁺X⁻ when X is an anion with poor coordinating capacity (ClO₄, BF₄) and the reaction is done in the presence of a neutral ligand (L₁ = PPh₃).

=> s l4 and silver

328333 SILVER

L10 36 L4 AND SILVER

=> s l10 and amidinate

326 AMIDINATE

L11 1 L10 AND AMIDINATE

=> d bib abs

L11 ANSWER 1 OF 1 CAPLUS COPYRIGHT 2007 ACS on STN
 AN 1999:757338 CAPLUS
 DN 132:72773
 TI Synthesis and Characterization of Functionalized N,N'-Diphenylformamidinate Silver(I) Dimers: Solid-State Structures and Solution Properties
 AU Archibald, Stephen J.; Alcock, Nathaniel W.; Busch, Daryle H.; Whitcomb, David R.
 CS Department of Chemistry, University of Kansas, Lawrence, KS, 66045, USA
 SO Inorganic Chemistry (1999), 38(24), 5571-5578
 CODEN: INOCAJ; ISSN: 0020-1669
 PB American Chemical Society
 DT Journal
 LA English
 AB Functionalized N,N'-diphenylformamidines and their deprotonated Ag(I) complexes were synthesized: Ag(I) N,N'-bis(4-alkylphenyl)formamidinate (alkyl = Me, Et, Bu, and n-hexadecyl, 1-4, resp.); Ag(I) N,N'-bis(4-trifluoromethylphenyl)formamidinate (5), Ag(I) N,N'-bis(3-methoxyphenyl)formamidinate (6), Ag(I) N,N'-bis(3-methylthiophenyl)formamidinate (7), Ag(I) N,N'-bis(2-methoxyphenyl)formamidinate (8), Ag(I) N,N'-bis(2-

methylthiophenyl)formamidinate (9). The effects of increasing the coordination number of the Ag(I) centers by donor substituents on the Ph groups were studied by solution and solid-state studies. Variable-temperature

1H

NMR (223-303 K) for 1-4 shows coupling between the proton attached to the amidinate C and the 107/109Ag centers at room temperature which is unaffected by cooling 2. For the four-coordinate complexes 8 and 9, such coupling is only observed on cooling. Mol. weight measurements recorded in solution by vapor pressure osmometry at 310 K show some aggregation to higher mol. weight species than simple dimers for 1-4 and 6, but 8 and 9 exist as discrete dimeric species. Measurement of thermal stability shows the expected increase in stability with increasing coordination number. Compds. 8 and 9 were structurally characterized by x-ray methods. Both show four-coordinate Ag dimers bridged by two amidinate ligands with addnl. longer interactions with the ether oxygens or thioether sulfurs.

RE.CNT 79 THERE ARE 79 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT